***Chemistry***

**13: Fundamental Equilibrium Concepts**

**13.1: Chemical Equilibria**

1. What does it mean to describe a reaction as “reversible”?

Solution

The reaction can proceed in both the forward and reverse directions.

3. If a reaction is reversible, when can it be said to have reached equilibrium?

Solution

When a system has reached equilibrium, no further changes in the reactant and product concentrations occur; the reactions continue to occur, but at equivalent rates.

5. If the concentrations of products and reactants are equal, is the system at equilibrium?

Solution

The concept of equilibrium does not imply equal concentrations, though it is possible.

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